



Kuwait University
Office of Assistant Vice President for Evaluation
and Measurement

Academic Aptitude Tests

Student Name

Version A

Civil ID No.

Instructions:

1. The aptitude tests consist of three tests.

<u>Test</u>	<u>Number of Questions</u>	<u>Time</u>
English	85	1 Hour
Mathematics	20 (No Calculator)	1 Hour
Chemistry	25	1 Hour

2. Mark all your answers on the **Answer Sheet** and in the proper section. On your answer sheet as shown below, using a pencil, darkenthe proper circle.



3. Verify all personal and test data on answer sheet and don't make any changes unless approved by the proctor.
4. Write down your name and Civil ID# on the test booklet.
5. Copy the test's version on your answer sheet.
6. Follow the proctor's instruction during the test.
7. During testing, be quite and avoid any cheating situation.
8. Observe the allocated and the announced time for each test.

Chemistry Test

Atomic Molar Mass (g/mol):

Nitrogen (N) = 14.0

Oxygen (O) = 16.0

Atomic Number:

Hydrogen (H) = 1

Carbon (C) = 6

Nitrogen (N) = 7

Oxygen (O) = 8

Sodium (Na) = 11

Chlorine (Cl) = 17

Physical Constants:

Ion product constant for water (K_w) at 25 °C = 1.00×10^{-14}

- Mixing olive oil with an aqueous solution of table salt (NaCl) will form:
 - homogeneous mixture
 - heterogeneous mixture
 - colloidal solution
 - suspension
- Which of the following statements is **true**?
 - Sulfur (S) burns in air to form sulfur dioxide (SO₂) gas
 - Iron metal (Fe) forms iron carbonate when exposed to air
 - Sodium hydroxide (NaOH) reacts with nitric acid (HNO₃) to form salt, water and gas
 - All basic solutions turn blue litmus paper to red
- In which of the following pairs, do both compounds exist as liquids at room temperature?
 - Nickel chloride (NiCl₂), and Bromine (Br₂)
 - Sodium acetate (CH₃COONa), and Nickel Chloride (NiCl₂)
 - Ethanol (C₂H₅OH), and Bromine (Br₂)
 - Sulfur dioxide (SO₂), and Ethanol (C₂H₅OH)
- Which of the following compounds is classified as salt?
 - HCN
 - CH₃COOH
 - Ca₃(PO₄)₂
 - N₂O₅
- Which of the following anions contains sulfur atom?
 - Carbonate
 - Dichromate
 - Phosphate
 - Thiocyanate
- When an atom gains two electrons it becomes:
 - Doubly negative charged
 - Triply positive charged
 - Doubly positive charged
 - Neutral
- How many ions would be produced on dissolving one formula unit of potassium hydrogen iodate (KH(IO₃)₂) in water?
 - 8
 - 10
 - 4
 - 3
- Which of the following represents a conjugate acid-base pair?
 - HBr(aq) and HCl(aq)
 - HCO₃⁻(aq) and CO₃²⁻(aq)
 - H₃O⁺(aq) and NH₄⁺(aq)
 - H₂O(l) and H₂O₂(aq)

25. A 5.00 cm³ of sulfuric acid (H₂SO₄) aqueous solution was titrated with 0.050 M of potassium hydroxide (KOH) standard solution. The titration required 7.10 cm³ of the base for complete neutralization. What is the acid's concentration?
- a. 0.0355 mole / liter c. 0.0178 mole / liter
b. 0.0500 mole / liter d. 0.0710 mole / liter

Answers - English Exam		إجابات اختبار اللغة الانجليزية		Answers - English Exam		إجابات اختبار اللغة الانجليزية		Answers - English Exam		إجابات اختبار اللغة الانجليزية	
Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers
1 -	A B C D	19 -	A B C D	37 -	A B C D	55 -	A B C D	73 -	A B C D		
2 -	A B C D	20 -	A B C D	38 -	A B C D	56 -	A B C D	74 -	A B C D		
3 -	A B C D	21 -	A B C D	39 -	A B C D	57 -	A B C D	75 -	A B C D		
4 -	A B C D	22 -	A B C D	40 -	A B C D	58 -	A B C D	76 -	A B C D		
5 -	A B C D	23 -	A B C D	41 -	A B C D	59 -	A B C D	77 -	A B C D		
6 -	A B C D	24 -	A B C D	42 -	A B C D	60 -	A B C D	78 -	A B C D		
7 -	A B C D	25 -	A B C D	43 -	A B C D	61 -	A B C D	79 -	A B C D		
8 -	A B C D	26 -	A B C D	44 -	A B C D	62 -	A B C D	80 -	A B C D		
9 -	A B C D	27 -	A B C D	45 -	A B C D	63 -	A B C D	81 -	A B C D		
10 -	A B C D	28 -	A B C D	46 -	A B C D	64 -	A B C D	82 -	A B C D		
11 -	A B C D	29 -	A B C D	47 -	A B C D	65 -	A B C D	83 -	A B C D		
12 -	A B C D	30 -	A B C D	48 -	A B C D	66 -	A B C D	84 -	A B C D		
13 -	A B C D	31 -	A B C D	49 -	A B C D	67 -	A B C D	85 -	A B C D		
14 -	A B C D	32 -	A B C D	50 -	A B C D	68 -	A B C D				
15 -	A B C D	33 -	A B C D	51 -	A B C D	69 -	A B C D				
16 -	A B C D	34 -	A B C D	52 -	A B C D	70 -	A B C D				
17 -	A B C D	35 -	A B C D	53 -	A B C D	71 -	A B C D				
18 -	A B C D	36 -	A B C D	54 -	A B C D	72 -	A B C D				

Answers - Mathematics Exam		إجابات اختبار الرياضيات		Answers - Mathematics Exam		إجابات اختبار الرياضيات	
Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers
1 -	A B C D	6 -	A B C D	11 -	A B C D	16 -	A B C D
2 -	A B C D	7 -	A B C D	12 -	A B C D	17 -	A B C D
3 -	A B C D	8 -	A B C D	13 -	A B C D	18 -	A B C D
4 -	A B C D	9 -	A B C D	14 -	A B C D	19 -	A B C D
5 -	A B C D	10 -	A B C D	15 -	A B C D	20 -	A B C D

Answers - Chemistry Exam		إجابات اختبار الكيمياء		Answers - Chemistry Exam		إجابات اختبار الكيمياء	
Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers
1 -	A B C D	6 -	A B C D	11 -	A B C D	16 -	A B C D
2 -	A B C D	7 -	A B C D	12 -	A B C D	17 -	A B C D
3 -	A B C D	8 -	A B C D	13 -	A B C D	18 -	A B C D
4 -	A B C D	9 -	A B C D	14 -	A B C D	19 -	A B C D
5 -	A B C D	10 -	A B C D	15 -	A B C D	20 -	A B C D
						21 -	A B C D
						22 -	A B C D
						23 -	A B C D
						24 -	A B C D
						25 -	A B C D

Answers - Arabic Exam		إجابات اختبار اللغة العربية		Answers - Arabic Exam		إجابات اختبار اللغة العربية		Answers - Arabic Exam		إجابات اختبار اللغة العربية	
Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers	Q's#	Answers
1 -	A B C D	11 -	A B C D	21 -	A B C D	31 -	A B C D	41 -	A B C D	51 -	A B C D
2 -	A B C D	12 -	A B C D	22 -	A B C D	32 -	A B C D	42 -	A B C D	52 -	A B C D
3 -	A B C D	13 -	A B C D	23 -	A B C D	33 -	A B C D	43 -	A B C D	53 -	A B C D
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6 -	A B C D	16 -	A B C D	26 -	A B C D	36 -	A B C D	46 -	A B C D	56 -	A B C D
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